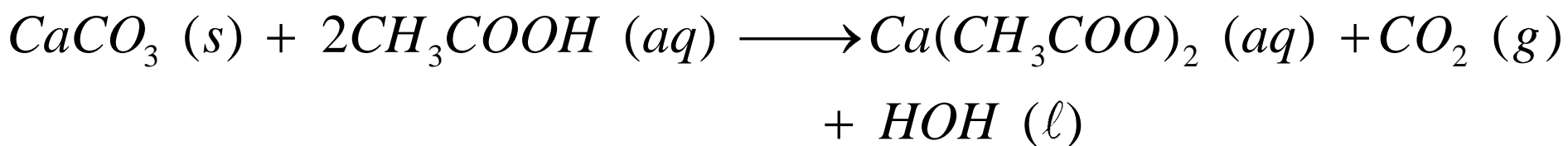
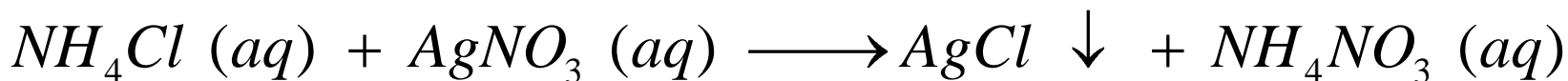
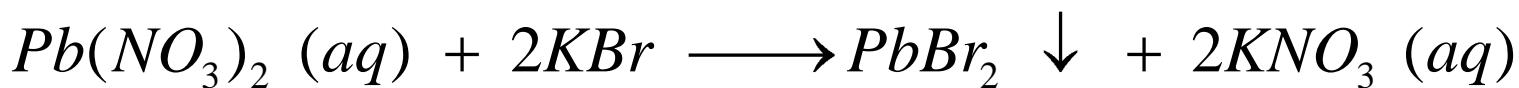
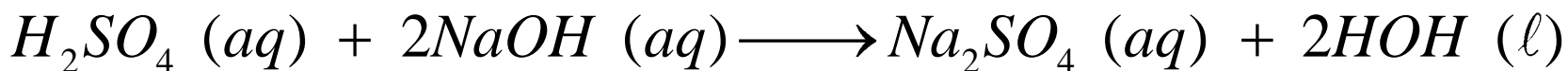
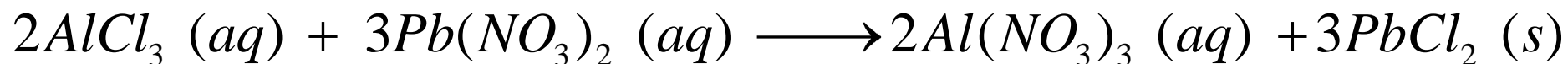
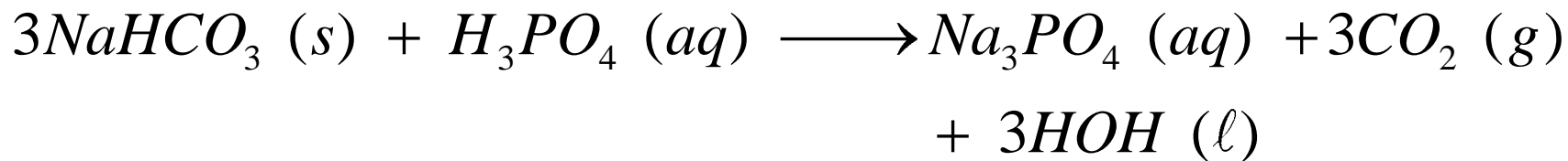
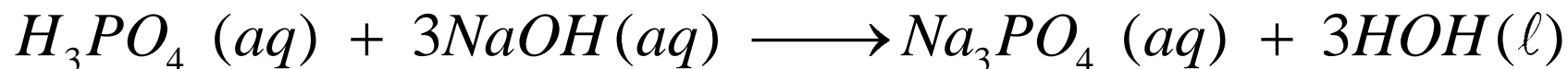
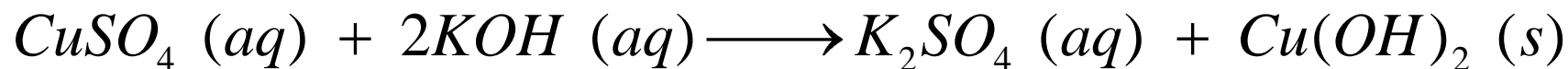


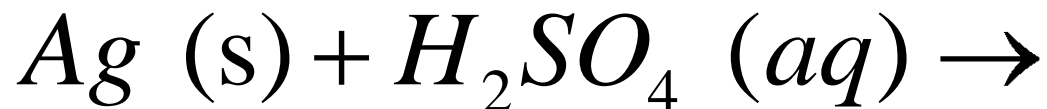
Pg. 173 #11



Pg. 173 #12



Pg. 173 #13a

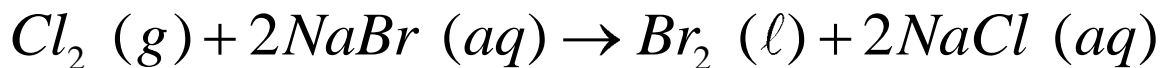
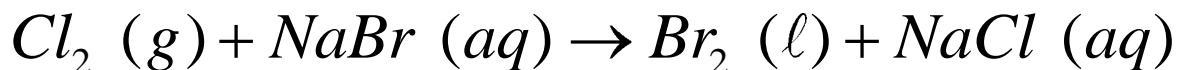


The reaction does not work because silver is **less reactive** than hydrogen. That is, it is lower on the activity series.

Pg. 173 #13b



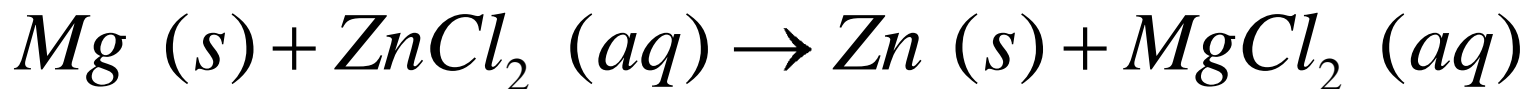
This reaction **will work** because chlorine is more reactive than bromine. That is, it is higher on the activity series.



Pg. 173 #13c



This reaction **will work** because magnesium is more reactive than zinc. That is, it is higher on the activity series.

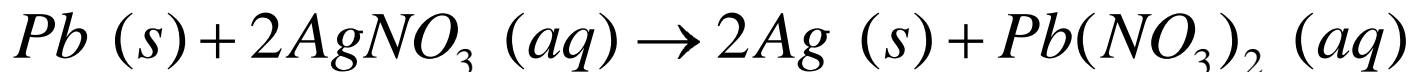
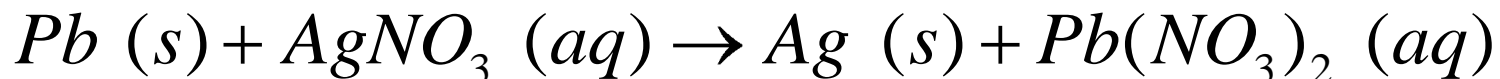


This reaction is already balanced.

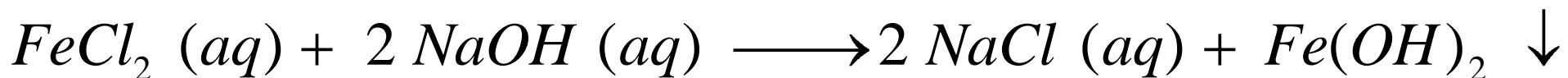
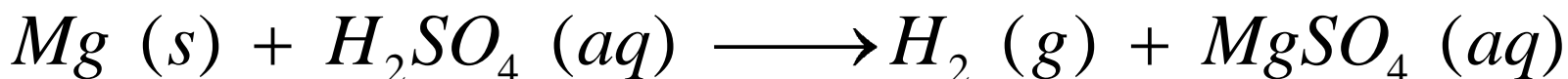
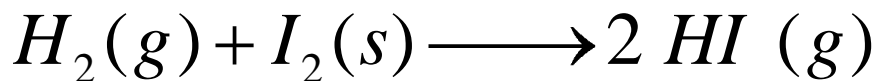
Pg. 173 #13d



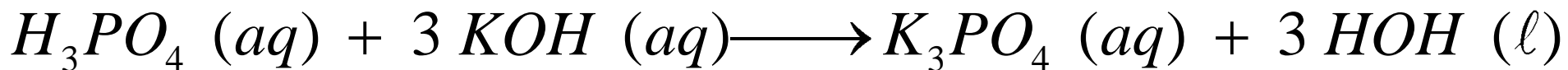
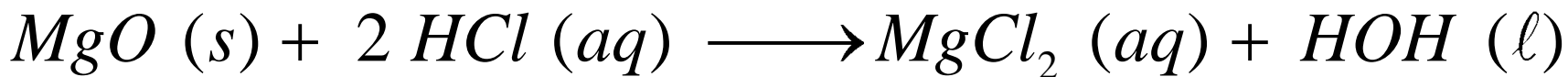
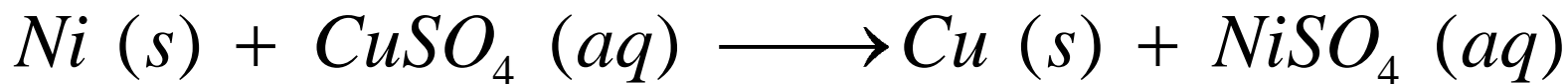
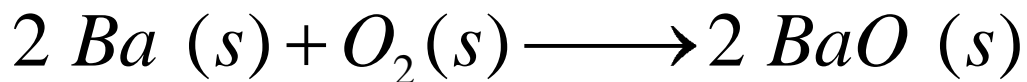
This reaction **will work** because lead is more reactive than silver. That is, it is higher on the activity series.



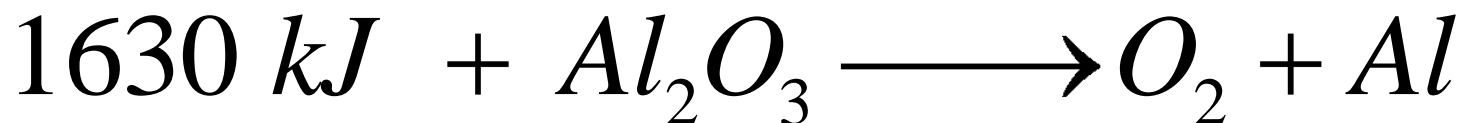
Pg. 173 #15



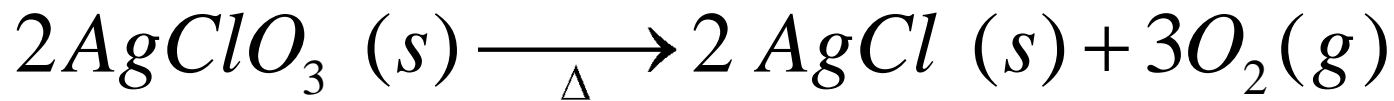
Pg. 173 #17



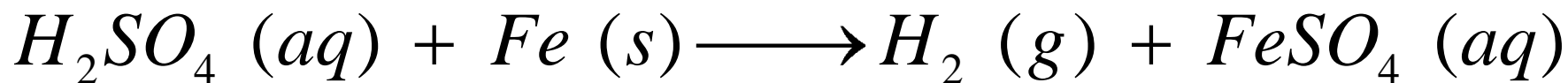
Pg. 173 #21



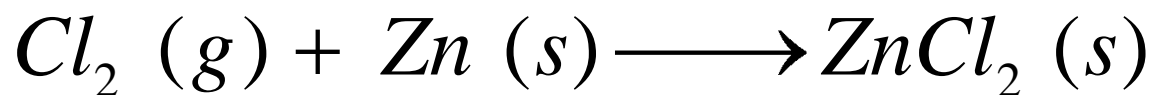
Pg. 174 #23



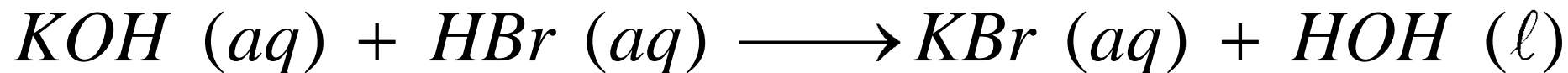
decomposition



single displacement

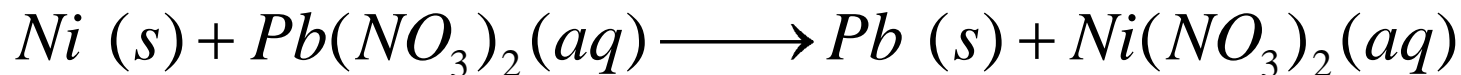


synthesis

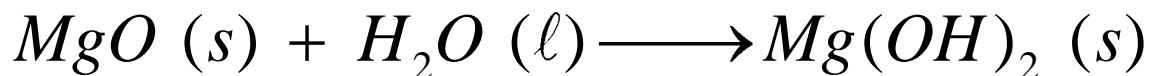


double displacement/acid-base neutralization

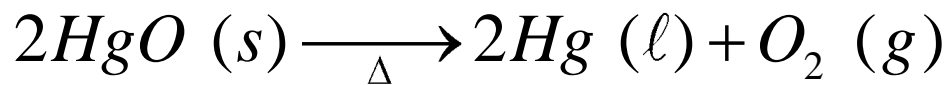
Pg. 174 #24



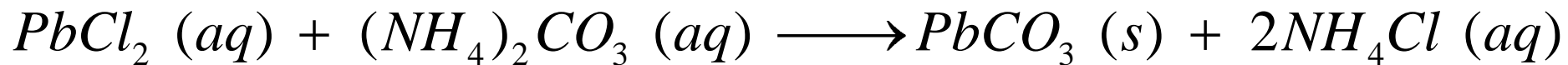
double displacement



synthesis



decomposition

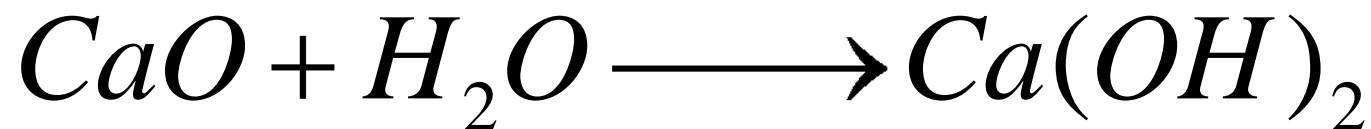
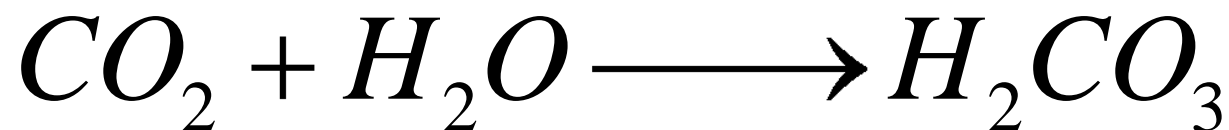
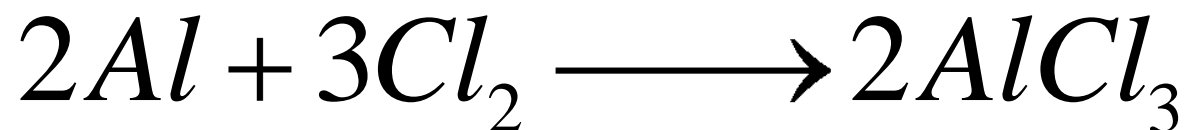
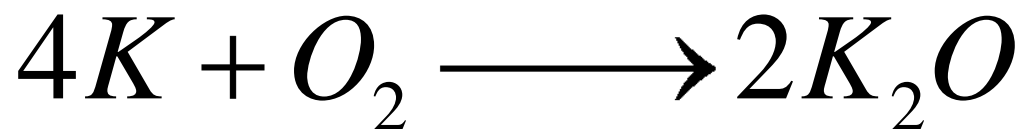


double displacement

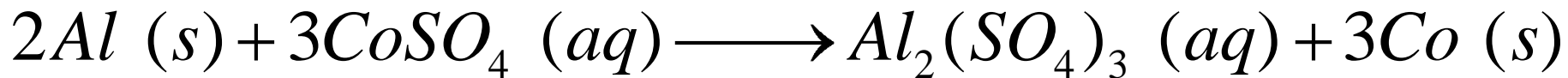
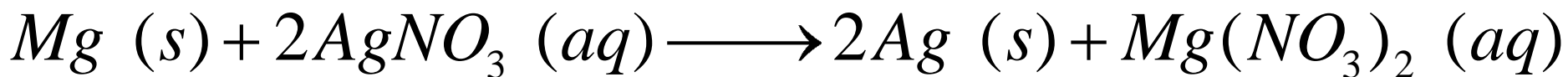
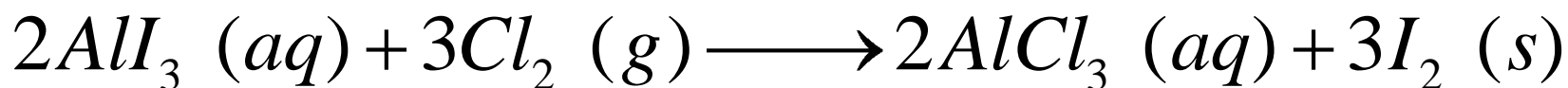
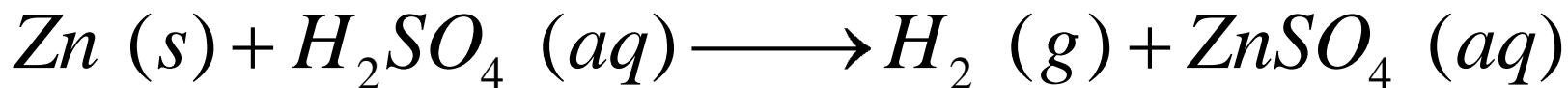
Pg. 174 #30

This reaction will not take place because zinc is less reactive than magnesium. For a single displacement reaction to work, the free state element must be more reactive than the element it is trying to replace.

Pg. 174 #32



Pg. 174 #34



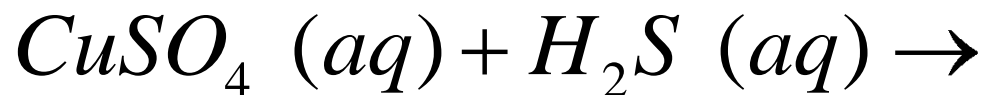
Pg. 174 #35a



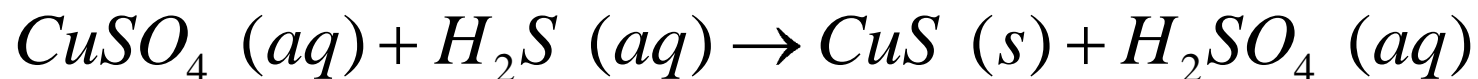
This reaction will work because zinc hydroxide will be a precipitate.



Pg. 174 #35b

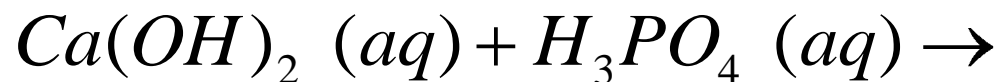


This reaction will work because copper(II) sulfide will be a precipitate.

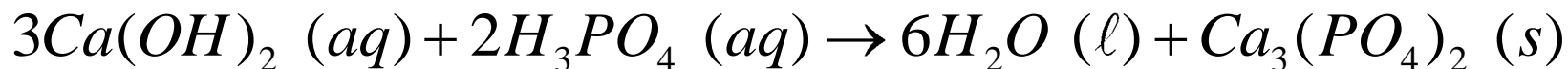


This reaction is already balanced.

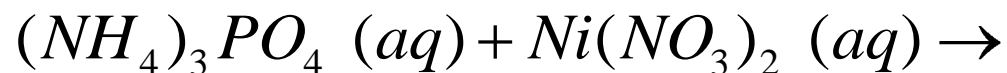
Pg. 174 #35c



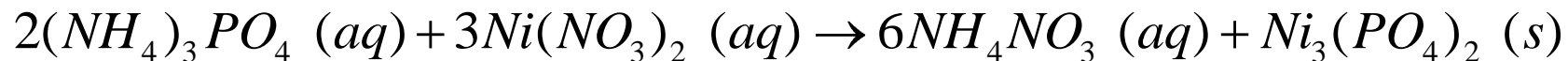
This reaction will work because calcium phosphate will be a precipitate and water is produced.



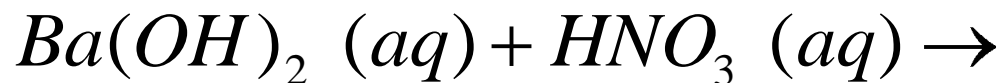
Pg. 174 #35d



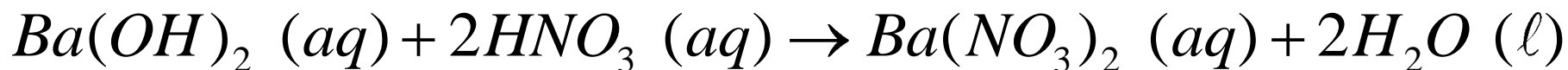
This reaction will work because nickel(II) phosphate will be a precipitate.



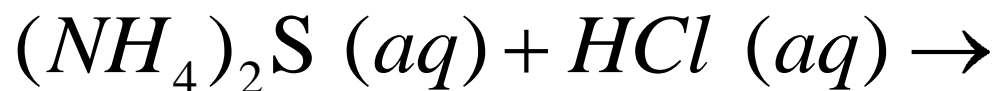
Pg. 174 #35e



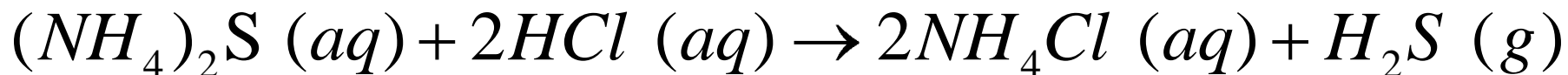
This reaction will work because water is produced.



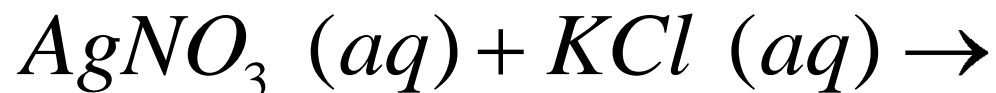
Pg. 174 #35f



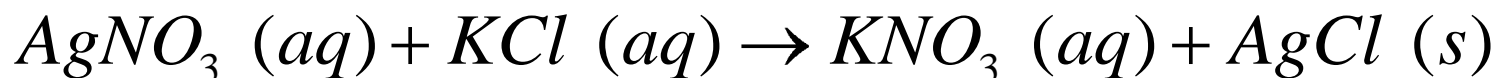
This reaction works because hydrogen sulfide is a gas.



Pg. 174 #36a

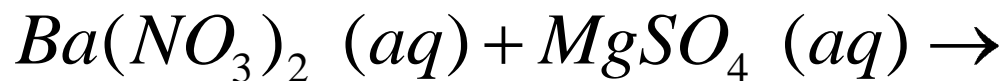


This reaction will work because silver chloride will be a precipitate.

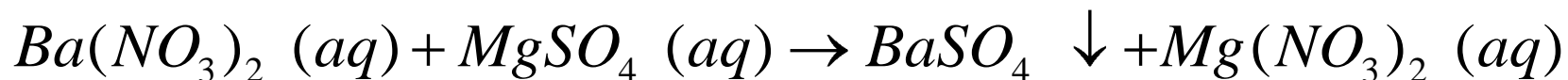


This reaction is already balanced.

Pg. 174 #36b



This reaction will work because barium sulfate will be a precipitate.

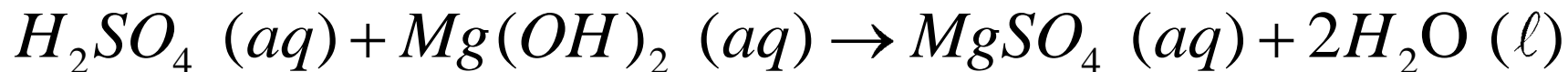


This reaction is already balanced.

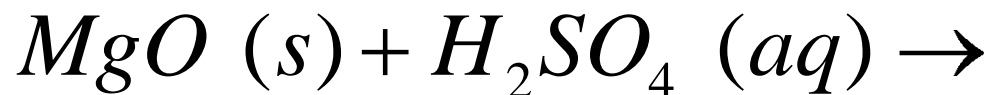
Pg. 174 #36c



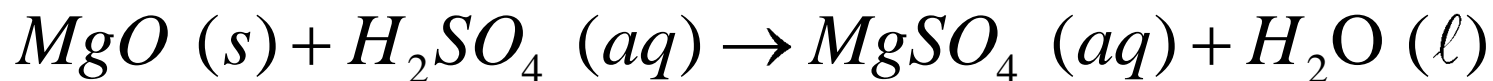
This reaction will work because the reaction produces water, which is not aqueous.



Pg. 174 #36d

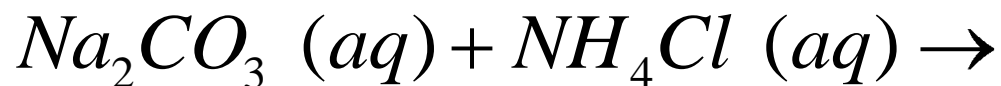


This reaction will work because the reaction produces water, which is not aqueous.



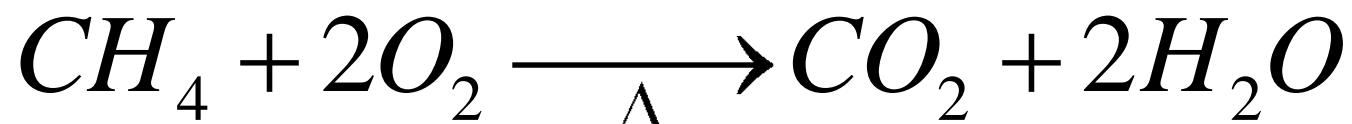
This reaction is already balanced.

Pg. 174 #36e



This reaction will not work because the predicted products will contain either sodium or ammonium, both of which make aqueous compounds.

Pg. 174 #37



Pg. 175 #38

